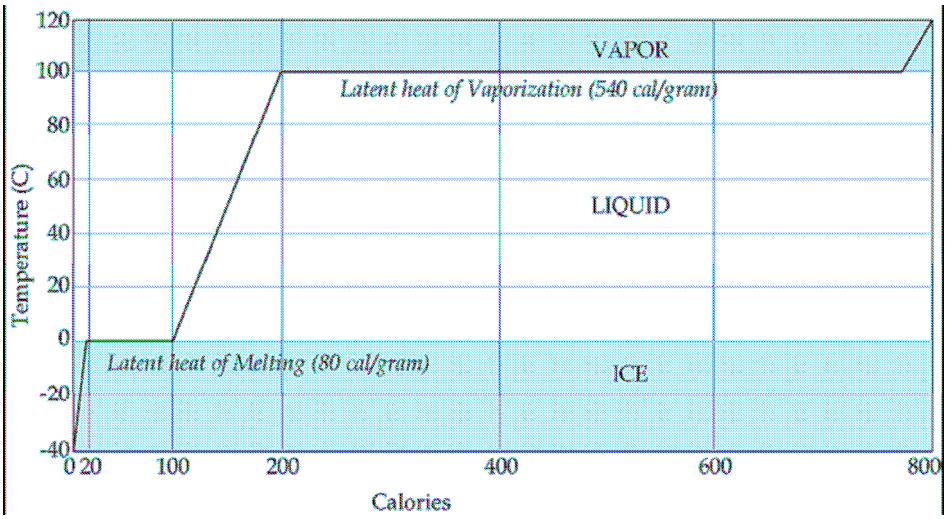

Specific heat of water: 1 calorie/gram ℃
Specific heat of water vapor: 0.44 cal/gram ℃
Specific heat of ice: 0.5 calorie/gram ℃
Latent heat of melting: 80 calorie/gram
Latent heat of vaporization: 540 calorie/gram

START: 1 gram of solid ice at -10℃
END: 1 gram of water vapor at 120℃

Temp change	State	Heat change
-10° to 0℃	S	***1gx10℃x0.5cal/g℃=5 cal***
0℃	S → L	***1gx80 cal/g=80 cal***
0° to 100℃	L	***1gx100℃x1cal/g℃ =100 cal***
100℃	L → G	***1gx540cal/g=540 cal***
100° to 120℃	G	***1gx20℃x0.44cal/g℃=8.8 cal***
	Total	***733.8 cal***

START: 5 gram of solid ice at -20℃
END: 5 gram of water vapor at 105℃

Temp change	State	Heat change
-20° to 0℃	S	
0℃	S → L	
0° to 100℃	L	
100℃	L → G	
100° to 105℃	G	
	Total	

START: 10 gram of water at 5℃
END: 5 gram of water vapor at 100℃

Temp change	State	Heat change
	S	
	S → L	
	L	
	L → G	
	G	
	Total	

START: 10 gram of Ice at -50℃
END: 10 gram of water vapor at 170℃

Temp change	State	Heat change
	S	
	S → L	
	L	
	L → G	
	G	
	Total	

START: 5 gram of solid ice at -10℃
END: How does it end after the addition of 5700 calories

Temp change	State	Heat change
	S	
	S → L	
	L	
	L → G	
	G	
	Total	